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Solve the following stoichiometry grams-grams problems:

6) Using the following equation:  $2 \text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$  How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?

7) Using the following equation:  $\text{Pb}(\text{SO}_4)_2 + 4 \text{LiNO}_3 \rightarrow \text{Pb}(\text{NO}_3)_4 + 2 \text{Li}_2\text{SO}_4$

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product (reverse the process for liters to grams):

- Use factor label method
- Use mass of reactant from the Periodic Table
- 1 mol = \_\_\_\_\_ g
- Use the mole to mole ratio from the balanced reaction

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